

## **Remarks**

Claims 5 - 11 are pending. Claims 7 and 8 have been amended. Favorable reconsideration is respectfully requested.

The claims are directed to a novel method of removing aluminum chloride from organochlorosilanes, preferably organochlorosilanes from the direct synthesis. Organochlorosilanes are silicon compounds containing at least one organo group such as a methyl, ethyl, propyl or phenyl group, and at least one chlorine, bound to silicon. The compounds may additionally contain silicon-bonded hydrogen. Thus, examples are  $\text{CH}_3\text{SiCl}_3$ ,  $(\text{CH}_3)_2\text{SiCl}_2$ ,  $(\text{CH}_3)_3\text{SiCl}$ ,  $(\text{CH}_3)_2\text{HSiCl}$ ,  $\text{CH}_3\text{H}_2\text{SiCl}$ ,  $\text{CH}_3\text{HSiCl}_2$ , and the like, exemplified here as the most common silanes where the organo group is a methyl group.

Claims 7, 8, and 11 have been rejected under 35 U.S.C. § 112, and claims 7 and 8 have been amended to recite that the number “a” of organogroups is 1, 2, or 3, and at least one chlorine is present, as is clear from the designation “organochlorosilane.” Applicants now believe the claims to be clear in this respect, and solicit withdrawal of the rejection of the claims under 35 U.S.C. § 112.

Claims 5 - 8 have been rejected under 35 U.S.C. § 102(a) over Japanese published application JP 2005-029428 (“*Wakamatsu*”) which was published on February 3, 2005. Applicants respectfully traverse this rejection as first, *Wakamatsu* is not a proper reference, as Applicants’ priority date, to which priority was claimed and is of record, is March 23, 2004. Enclosed herewith is an English translation of the German priority document, which Applicants’ attorney states is an accurate translation. Withdrawal of *Wakamatsu* as a reference is respectfully solicited.

Second, *Wakamatsu*, is not directed to the claimed subject matter. *Wakamatsu* is concerned with the purification of chlorosilanes such as tetrachlorosilane, not organochlorosilanes. Nor are *Wakamatsu*’s chlorosilanes prepared by the direct synthesis. In

the field of silane production, the “direct process” has an art recognized meaning: preparing of organohalosilanes from organohalides by reaction with metallic silicon, also termed the Müller-Rochow process. While reaction of chlorine with silicon to form tetrachlorosilane may be a “direct synthesis from the elements, it is not the “direct synthesis” in art terms. Copies of relevant pages from Walter Noll, CHEMISTRY AND TECHNOLOGY OF SILICONES, Academic Press, Orlando, Florida, ©1968 are enclosed.

As *Wakamatsu* is not a proper reference, and as this reference is also not relevant to the claimed subject matter, withdrawal of the rejection of the claims under 35 U.S.C. § 102(a) over *Wakamatsu* is solicited.

Claims 5 - 11 have been rejected under 35 U.S.C. § 103(a) over *Wakamatsu* in view of Geisberger U.S. 5,434,286 (“*Geisberger*”). *Geisberger* is not directed to removal of AlCl<sub>3</sub> from organochlorosilanes, but rather is directed to redistribution of dimethyldichlorosilane to form dimethylchlorosilane [(CH<sub>3</sub>)<sub>2</sub>HSiCl] in the presence of a heterogenous catalyst. *Geisberger* therefore does not appear relevant to the claimed subject matter. However, as *Wakamatsu* is not a proper reference, withdrawal of this rejection under 35 U.S.C. § 103(a) over the combination of *Wakamatsu* and *Geisberger* is also respectfully solicited.

Applicants submit that the claims are now in condition for Allowance, and respectfully request a Notice to that effect. If the Examiner believes that further discussion will advance the prosecution of the Application, the Examiner is highly encouraged to telephone Applicants' attorney at the number given below.

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Respectfully submitted,

**Wilfried Kalchauer et al.**

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Attachments: pages from Noll  
copy of English translation of PCT

# **Chemistry and Technology of Silicones**

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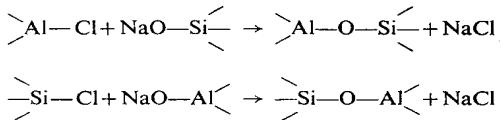
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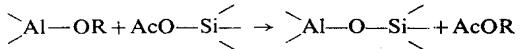
### 7.1.2 ALUMOSILOXANES

In addition to cleavage reactions of the Si—O—Si bond with aluminum halides (Section 5.5.5) the following routes are available for synthesis of alumosiloxanes.

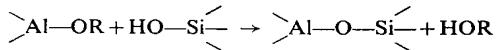
(a) The reaction of aluminum chlorides (19, 252) or organosiloxyaluminum chlorides (156) with alkali-metal silanolates or, conversely, of chlorosilanes with sodium aluminate (27, 281) proceeds very smoothly:



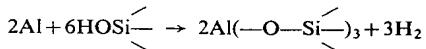
(b) The reaction of aluminum alkoxides with acyloxy silanes (168, 186, 244):



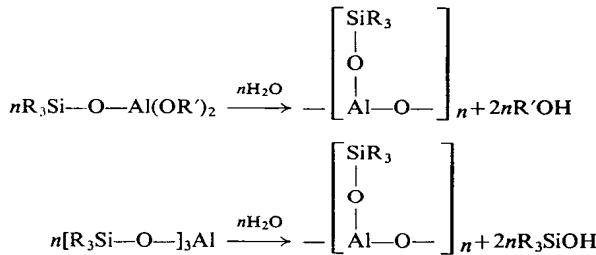
(c) The reaction of aluminum alkoxides with silanols, e.g., of aluminum butoxide with triethylsilanol (213):



(d) Finally, the reaction of aluminum with silanols (11, 308):

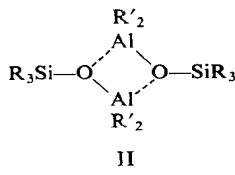
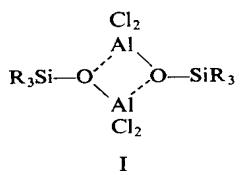


Attempts have been made to obtain high-molecular-weight products not only directly by reactions of this type (i.e., using polyfunctional starting materials) but also by a stepwise preparation of definite low-molecular-weight alumosiloxanes and subsequent partial hydrolysis (20, 151, 315):



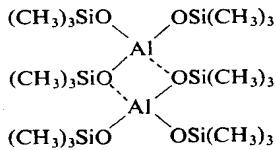
Whether cohydrolysis of aluminum chloride with organochlorosilanes (11) leads to polymers with a statistically homogeneous distribution of Al—O and Si—O units must remain very questionable, since the differences in the sensitivity of the aluminum and silicon halides to hydrolysis appears to be very great.

A number of definite low-molecular-weight aluminosiloxanes have been prepared and thoroughly investigated (253). The organosiloxyaluminum halides obtained by the cleavage of hexaorganodisiloxanes with aluminum halides are dimeric by coordination and contain the almost planar skeleton of structure (I). By alkylation with alkylolithiums these have been converted into the corresponding pentaalkylalumosiloxane compounds, which are also dimeric (II). These are thermally very stable and decompose only above 180°C, with the liberation of tetraalkylsilanes. They are sensitive to hydrolysis and oxidation but are less reactive than alkylaluminum compounds (253):

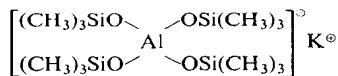


The hydrides obtained from (I) by means of lithium tetrahydroaluminate tend to polymerize and are stable in their dimeric form for only short periods (253).

The tris-triorganosiloxy compounds of aluminum have also been shown to form coordination dimers (253):



They are remarkably stable compounds which decompose only slowly above 280°C to siloxanes and alumina (253). They are relatively insensitive to water alone but are rapidly hydrolyzed in the presence of organic solvents (253). With alkali-metal silanolates, e.g.,  $(\text{CH}_3)_3\text{SiOK}$ , they form alkali-metal tetrakis-trimethylsilyl-aluminates smoothly and completely even at room temperature (253):



Linear, branched, and cross-linked polymers can be formed in the synthesis of high-molecular-weight aluminosiloxanes, depending on the nature of the starting material and on the reaction mechanism, but their structures must still be considered uncertain.

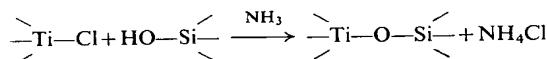
Cross-linked polymers form resinous or glasslike materials, which are said to be useful for the preparation of laminates (316) or for bonding glass fibers and phenol-formaldehyde resins (15). There is, however, no convincing evidence that aluminosiloxanes show decisive advantages over the simple siloxanes, if the complete pattern of their properties is evaluated. It has not been possible to obtain products analogous to silicone rubber (287), presumably because linearity cannot be achieved at high molecular weights because of unavoidable cross-linking through coordinative or primary valences.

### 7.1.3 TITANOSILOXANES

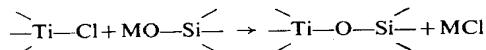
Compounds with Ti—O—Si bonds are formed when siloxanes are cleaved with titanium tetrachloride or butyl titanate (Section 5.5.5).

However, it is preferred to use the following processes to prepare titanatosiloxanes:

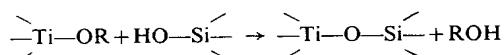
(a) The reaction of titanium tetrachloride with silanols in the presence of ammonia as an acid acceptor (90):



(b) The reaction of titanium tetrachloride with alkali-metal silanolates (9, 21, 107):

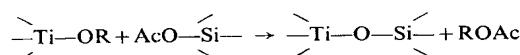


(c) The reaction of alkyl titanates with silanols (6, 38, 75, 80, 137):



This type of transesterification with alkyl titanates is catalyzed by sodium (80).

(d) The reaction of titanic acid esters and acetoxy silanes (108, 245):



As with the analogous reactions of boric acid esters and acetoxy silanes, there is again obviously the possibility of partial exchange reactions of —OR and —OAc groups which result in *intramolecular* rather than *intermolecular* reactions (13). Mono-, bis-, tris-, and tetrakis-trimethylsilyl titanates can, however, be obtained in good yield by this reaction if the stoichiometry of the components is carefully controlled (245).

Method for separating aluminium chloride from organochlorosilanes

The invention relates to a process for removing AlCl<sub>3</sub> 5 from a compound mixture comprising organochlorosilanes.

In the preparation of organochlorosilanes, product streams are obtained as a function of the process/operation and may comprise aluminium chloride 10 in different concentrations. Depending on the conditions, for example temperature, chemistry of the organosilanes, the AlCl<sub>3</sub> is dissolved at least partly in the liquid organochlorosilanes stream and thus cannot be removed by filtration. Being a Lewis acid, 15 AlCl<sub>3</sub> can exert highly disruptive influences depending on the process temperature.

An effective distillative removal of organochlorosilanes and AlCl<sub>3</sub> is possible only at 20 moderate temperatures up to about 150°C, since AlCl<sub>3</sub> exhibits the tendency to sublime and at least partly distils overhead with the organochlorosilanes when higher temperatures are employed.

According to Ullmann's Encyclopaedia of Industrial Chemistry, AlCl<sub>3</sub> has a sublimation temperature of 181.2°C at 101.3 kPa. 25

Disruptive influences of AlCl<sub>3</sub> are, for example:

30

1. At relatively high temperatures, AlCl<sub>3</sub> reacts with siloxanes, for example methylchlorodisiloxanes, to form aluminosiloxanes. Depending on the degree of branching

of these aluminosiloxanes, they are viscous to solid and can thus displace plant parts such as pipelines, or greatly reduce the heat transfer as a result of deposits in the region of heated plant parts such as  
5 heat exchangers.

The formation of such aluminosiloxanes is described, for example, in "W. Noll, Chemistry and Technology of Silicones, 1968, pages 238 and 340-342 (1).

10 2.  $\text{AlCl}_3$ , even at low temperatures, is an excellent catalyst for the exchange of the organic ligands and the Cl and H ligands in organochlorosilanes, especially in the presence of Si-H compounds. These reactions are sometimes used selectively for the preparation of  
15 certain organochlorosilanes. However, this ligand exchange may also be disruptive when the  $\text{AlCl}_3$  has not been added deliberately and the product spectrum is thus shifted in an undesired direction.

Such ligand exchange reactions are described, for  
20 example, in (1), pages 57-66 and J. Organomet. Chem. 260 (1984), 31-39, H. Schmöller, E. Hengge (2).

FR 2761360 states that, for example, the selective addition of compounds of the (R)-Si-(OR) type can  
25 reduce the catalytic effectiveness of  $\text{AlCl}_3$  in relation to ligand exchange. However, this method has the following disadvantages:

- an additional substance has to be used, which causes  
30 costs, and
- subsequently has to be destroyed/disposed of again and
- complicates the distillative workup of the organochlorosilanes.

3.  $\text{AlCl}_3$  catalyses the decomposition of methylchlorodisilanes in the direction of methylchloromonosilanes and oligo-/polysilanes; in the presence of Si-H bonds, this reaction begins at as low as from approx. 105°C. The oligo-/polysilanes formed may be viscous to solid and insoluble, and may induce the problems described under point 1. The formation of these oligo-/polysilanes is described, for example, in (2).

The presence is particularly disruptive in the processes below:

1. Direct synthesis of methylchlorosilanes according to Müller-Rochow.

In the direct synthesis of methylchlorosilanes, silicon is reacted with  $\text{MeCl}$  in the presence of various catalysts at about 265-310°C. This forms a mixture of various methylchloro(hydro)silanes, and also methylchlorodisilanes, methylchlorodisiloxanes and hydrocarbons.

The Si used typically contains 0.1-0.3% Al, and the additional addition of Al, for example as an alloy, to the reaction system is likewise known. Irrespective of the source and of the form used,  $\text{AlCl}_3$  forms at least partly from the aluminium and, owing to the temperatures, the system comprising the reaction products and unconverted starting materials leaves via the gas phase.

30

In "Catalysed Direct Reactions of Silicon; K.M. Lewis, D.G. Rethwisch; Elsevier 1993; Chapter 1" (3), Figure 3 on page 18 shows a schematic flow diagram of the process, in which the solid-containing reaction

products from the direct synthesis are condensed, the solids are removed and the crude silane is fed to the distillation.

5 "Ullmann's Encyclopaedia of Industrial Chemistry Vol. A  
24, page 26" describes a similar process.

The liquid crude silane mixture prepared in this way comprises, in addition to the methylchloromonosilanes, also AlCl<sub>3</sub>, methylchlorodisilanes, disiloxanes and hydrocarbons. This means that the reactions described under "disruptive influences" occur during the distillative workup even when only the filtered crude silane mixture freed of solid is used further.

15 In (3) on page 22-28, the following further workup method is specified as an alternative:

"Gases separated in the cyclone and filter are fed into the bottom of a scrubber in which products with normal boiling points less than about 170°C are separated from metal chlorides and other higher boilers. The distillate is fractionated into an overhead stream containing compounds boiling at or below 71°C and a side stream composed essentially of the cleavable disilanes. The bottoms, containing solids and 25 methylchlorosilanes, are purged periodically and sent to waste disposal."

The disadvantages in this process are:

30 - Since the methylchlorosilanes are a mixture of many different substances having a wide boiling point range, it is not possible simultaneously to drive all utilizable products out of the bottoms of the scrubber and/or of the fractionation unit and to keep the

temperature for the driving-out of the organochlorosilanes so low that the disadvantages described do not occur.

In other words, the scrubber or the fractionation unit  
5 is operated at temperatures at which the reactions catalysed by AlCl<sub>3</sub> do not occur to a noticeable extent, and the loss of utilizable methylchloro(di)silanes is automatically accepted.

However, when these plants are operated at a higher  
10 temperature at which almost all utilizable products are driven out, the undesired side reactions occur to an increased extent, and the higher-boiling fractions, for example the disilane fraction, simultaneously comprise not inconsiderable proportions of entrained AlCl<sub>3</sub>.

15

- The residues which occur are suspensions composed of liquid organochlorosilanes and solids. A workup or disposal of such product streams is generally to be classified as problematic.

20

When Figures 4 and 5, page 25, 26 and illustrative text in (3) are considered, it can be seen that the AlCl<sub>3</sub> introduced with the crude MCS direct reaction mixture is discharged with the disilanes in the direction of column A and will cause the problems already described many times, in this region or in the downstream 25 disilane workup at the latest.

## 2. AlCl<sub>3</sub>-catalysed high boiler workup

30 EP 829484 A, for example, describes the AlCl<sub>3</sub>-catalysed cleavage of the high boilers from the methylchlorosilane synthesis by means of HCl or H<sub>2</sub> or corresponding mixtures. EP 155626 A, for example,

describes the AlCl<sub>3</sub>-catalysed conversion of high boilers and low boilers in the direction of better utilizable monosilanes.

In the workup of these reaction products, comparable  
5 problems occur to those which have been described for the direct synthesis.

### 3. Amine-catalysed disilane cleavage

Various methylchlorodisilanes which are obtained as a  
10 by-product in direct synthesis may be converted using hydrogen chloride directly to methylchloromonosilanes (disilane cleavage). This reaction is catalysed, for example, by tertiary amines such as tributylamine, and is described in (3) on page 30-31. However, AlCl<sub>3</sub> forms  
15 complexes with amines which have only a greatly reduced, if any, catalytic activity, i.e. when AlCl<sub>3</sub> is present in sufficient amounts in the disilane cleavage, the reaction comes to a standstill.

20 It is an object of the invention to provide a process which enables the AlCl<sub>3</sub> and any further solids to be removed from AlCl<sub>3</sub>-containing organochlorosilane streams in a simple manner even when a conventional distillative separation is not effective owing to the  
25 boiling point of the organochlorosilanes.

The invention provides a process for removing AlCl<sub>3</sub> from a compound mixture (C1) comprising organochlorosilanes and having an AlCl<sub>3</sub> content of  
30 > 200 ppm based on the content of organochlorosilanes, in which the compound mixture (C1) is diluted with compounds (C2) which are selected from organochlorosilanes or mixtures of chloromethane and

organochlorosilanes in such a way that a product stream (P) with < 15% solids at a simultaneous concentration of < 25% of components having a boiling point > 71°C at 1013 hPa is obtained, and this product stream (P) is  
5 separated in an evaporator unit at a temperature < 165°C into volatile compounds (VC) and AlCl<sub>3</sub>-containing solid (S), all concentration data being based on the weight.

10 By virtue of the dilution of the compound mixture (C1) with the substantially solids- and high boiler-free compounds (C2), as a result of the "entrainment effect", even organochlorosilanes having a boiling point which is above the evaporator temperature are  
15 virtually fully evaporated, so that the AlCl<sub>3</sub>-containing solid (S) obtained is dry and can thus be handled in a simple manner.

20 The compound mixture (C1) stems preferably from the workup of the product mixture from the direct synthesis of alkylchlorosilanes, or from the AlCl<sub>3</sub>-catalysed high boiler cleavage of the by-products of the direct synthesis.

25 In the direct synthesis of alkylchlorosilanes of the general formula (I), R<sub>a</sub>H<sub>b</sub>SiCl<sub>4-a-b</sub>, in which a is 1, 2, 3 or 4 and b is 0, 1 or 2, from silicon metal and alkyl chlorides R-Cl where R is an alkyl radical, the by-products formed are di- and oligosilanes, carbosilanes,  
30 siloxanes and high-boiling cracking products.

The compound mixture (C1) comprising organochlorosilanes has an AlCl<sub>3</sub> content of preferably

> 300 ppm, more preferably > 500 ppm. In addition to AlCl<sub>3</sub>, the compound mixture (C1) may comprise further substances in dissolved or undissolved form.

5 The concentration of organosilanes having a boiling point of about 71°C in the compound mixture (C1) is preferably > 25%, more preferably > 35%, in particular > 45%.

10 The compound mixture (C2) has an AlCl<sub>3</sub> content of preferably < 50 ppm, more preferably < 30 ppm, in particular < 20 ppm. The concentration of organosilanes having a boiling point of above 71°C in the compound mixture (C2) is preferably < 5%, more preferably < 3%, in particular < 1%. The compound mixture (C2) may 15 comprise further volatile components, for example chloromethane.

The mixing ratio of (C1) and (C2) is adjusted such that the solids content of the resulting product stream (P) 20 is < 15% by weight, preferably < 8% by weight, more preferably < 4% by weight, and the content of components having a boiling point above 71°C is < 25%, preferably < 15%, more preferably < 10%. If desired, it is possible to add, in addition to the compound 25 mixtures (C1) and (C2), further solids-containing streams, for example silicon dusts from the direct synthesis which are no longer utilizable or suspensions of Si dusts from the direct synthesis, as are obtained, for example, in cyclones, filters and sludges, with the 30 proviso that the abovementioned features for the product stream (P) are maintained. On this subject, Figure 3 in (3) is incorporated by reference.

The product stream (P) is separated preferably at

< 155°C, more preferably at < 145°C. The evaporator unit is configured in such a way that the separation is effected very rapidly into substantially dry solids (F) and gaseous volatile compounds (VC). Those skilled in  
5 the art also refer to such a separation as a flash evaporation, i.e. there is no still/bottom in which the product stream (P) is heated continuously. Such an evaporator unit may, for example, be a unit for spray evaporation, thin-layer or thin-film evaporation.

10

If desired, it is possible in the course of the rapid evaporation for further volatile or gaseous compounds to be metered in in such a way that the removal of the evaporated volatile compounds (VC) is favoured.  
15 However, these compounds have to behave in a chemically inert manner in the system. The gaseous components used may, for example, be noble gases or nitrogen; the volatile compounds used may, for example, be chloromethane.

20

The volatile compounds (VC) drawn off in gaseous form are substantially free of undesired AlCl<sub>3</sub> impurities; solids entrained in the evaporation may, if desired, once more be removed by means of filtration before or  
25 after the condensation.

All temperatures reported relate to atmospheric pressure (1013 hPa). The process according to the invention may also be employed at reduced and at  
30 elevated pressure. Since the sublimation temperatures and most of the boiling temperatures depend greatly on the pressure, the temperatures in these cases change correspondingly.

All concentration data reported are based on the weight.

### Examples

5

#### Apparatus:

A three-necked glass flask with attached dropping funnel, gas inlet and heat-insulated gas outlet is immersed into an electrically heated oil bath with the temperature [T]. If required, nitrogen may be introduced into the system via the gas inlet; the product steam [P] is metered in at such a rate that no noticeable fractions of liquid products are present in the heated flask in the steady state. The evaporated organochlorosilanes, hydrocarbons, chlorohydrocarbons and any nitrogen fed in leave the flask via the insulated gas outlet and are fed to a condensation system cooled with dry ice [D = condensate]. The solids introduced with the product stream [P] and most of the metal chlorides dissolved in [P] remain substantially in the flask.

The organochlorosilanes, hydrocarbons, chlorohydrocarbons were analysed by means of GC. The aluminium contents in the organochlorosilanes were determined, and the solids were analysed, by means of ICP.

All examples were carried out at standard pressure.

#### Comparative Example 1

30 The intention was to prove that a conventional distillation cannot effectively remove high-boiling organochlorosilanes and AlCl<sub>3</sub>.

In a typical laboratory batchwise distillation system, AlCl<sub>3</sub> and solid-containing crude silane mixture are

subjected to a fractional distillation up to a bottom temperature of 167°C. The resulting solids-free distillates 1-A to 1-E and the solids-containing liquid residue 1-R were analysed. The results are listed in  
 5 Table 1.

Table 1

Fraction	1-A	1-B	1-C	1-D	1-E	1-R
Top temperature (°C)	20-65	65-70	70-75	75-78	78-160	
Sum of dimethyltetra-chlorodisilane and trimethyltrichlorodisilane	3.5	3.7	8.2	12.3	78.8	83.7
Compounds having a boiling point > 75°C, without dimethyltetra-chlorodisilane and trimethyltrichlorodisilane	1.3	1.3	4.4	3.4	11.3	15.9
Aluminium content (ppm)	<3	<3	3	4	185	8200

## 10 Examples 2-4

The product streams [C1] and [C2] and 3% by weight of ultrafine silicon dusts, as described in (3), Figure 3, "Filter" were mixed in such a ratio that the product stream [P] was formed. This mixture was metered into  
 15 the apparatus described at an oil bath temperature of 140°C, 150°C and 160°C, and a gentle nitrogen stream was additionally fed in via the gas inlet. In all cases, the thus obtained solids-free condensates [D] contained an Al content of < 2 ppm. The results are  
 20 listed in Table 2.

Table 2

	C-1	C-2	P
Chloromethane	0	5.9	4.9
Dimethylchlorosilane	0	0.7	0.5
Methyldichlorosilane	0	8.2	7.3

Trimethylchlorosilane	0.2	3.8	3.3
Methyltrichlorosilane	0.8	7.3	6.6
Dimethyldichlorosilane	41.9	72.0	70.1
Sum of dimethyltetrachlorodisilane and trimethyltrichlorodisilane	47.7	0.0	4.8
Compounds having a boiling point > 71°C, without dimethyltetrachlorodisilane and trimethyltrichlorodisilane	9.4	2.1	2.6
Aluminium content (ppm)	750	5	80

## Example 5

Analogous to Example 2-4, with the alteration that the mixing ratio of [C1] and [C2] was altered and the product stream [P] thus had the composition below; the oil bath temperature was 160°C. The thus obtained solids-free condensate [D] contained an Al content of 6 ppm. The results are listed in Table 3.

10 Table 3

	P
Chloromethane	4.4
Dimethylchlorosilane	0.5
Methyldichlorosilane	6.5
Trimethylchlorosilane	3.1
Methyltrichlorosilane	6.0
Dimethyldichlorosilane	66.2
Sum of dimethyltetrachlorodisilane and trimethyltrichlorodisilane	9.6
Compounds having a boiling point > 71°C, without dimethyltetrachlorodisilane and trimethyltrichlorodisilane	3.6
Aluminium content (ppm)	170

## Example 6:

Analogous to Example 2-4, with the alteration that a different compound mixture [C2] was used, the additional metering of Si dusts was dispensed with and no nitrogen was metered into the system during the experiment. The thus obtained solids-free condensate [D] contained an Al content of < 2 ppm. The results are listed in Table 4.

Table 4

	C-1	C-2	P
Dimethylchlorosilane	0	0.2	0.2
Methyldichlorosilane	0	3.0	2.7
Trimethylchlorosilane	0.2	2.1	1.9
Methyltrichlorosilane	0.8	6.0	5.5
Dimethyldichlorosilane	41.9	87.8	83.2
Sum of dimethyltetrachlorodisilane and trimethyltrichlorodisilane	47.7	0.0	4.8
Compounds having a boiling point > 71°C, without dimethyltetrachlorodisilane and trimethyltrichlorodisilane	9.4	1.0	1.8
Aluminium content (ppm)	750	< 5	70

## Claims:

1. Process for removing AlCl<sub>3</sub> from a compound mixture (C1) comprising organochlorosilanes and having an AlCl<sub>3</sub> content of > 200 ppm based on the content of organochlorosilanes, in which the compound mixture (C1) is diluted with compounds (C2) which are selected from organochlorosilanes or mixtures of chloromethane and organochlorosilanes in such a way that a product stream (P) with < 15% solids at a simultaneous concentration of < 25% of components having a boiling point > 71°C at 1013 hPa is obtained, and this product stream (P) is separated in an evaporator unit at a temperature < 165°C into volatile compounds (VC) and AlCl<sub>3</sub>-containing solid (S), all concentration data being based on the weight.
2. Process according to Claim 1, in which the compound mixture (C1) comprising organochlorosilanes stems from the direct synthesis of alkylchlorosilanes or from the AlCl<sub>3</sub>-catalysed high boiler cleavage of the by-products of the direct synthesis.
- 25 3. Process according to Claim 1 or 2, in which the organochlorosilanes are alkylchlorosilanes of the general formula R<sub>a</sub>H<sub>b</sub>SiCl<sub>4-a-b</sub> in which a is 1, 2, 3 or 4, b is 0, 1 or 2, and R is a methyl, ethyl, butyl or propyl radical.

4. Process according to Claims 1 to 3, in which the evaporator unit used is a unit for spray evaporation, thin-layer or thin-film evaporation.